Correction to A Toolkit to Fit Nonbonded Parameters from and for Condensed Phase Simulations

Florent Hédin,[†] Krystel El Hage,[†] and Markus Meuwly^{*,†,‡}

[†]Department of Chemistry, University of Basel, Klingelbergstrasse 80, CH-4056 Basel, Switzerland

[‡]Department of Chemistry, Brown University, Providence, Rhode Island 02912, United States

J. Chem. Inf. Model. 2016, 56 (8), 1479–1489. DOI: 10.1021/acs.jcim.6b00280

I n a recent publication we presented a fitting environment for parametrizing point charge (PC) and multipolar (MTP) force fields for condensed-phase simulations.¹ After publication of this work it came to our attention that one of the scripts contained an error which caused an energy component in the free energy simulations to return incorrect values. This affects the optimization of the parameter *l* when scaling the Lennard-Jones (LJ) parameters according to $\varepsilon^* = l\varepsilon$ and $R^*_{\min}/2 = lR_{\min}/2$ but not the MTP terms.

Hence, all compounds considered were reparametrized according to the procedure described in ref 1. The corresponding correlations between experiment and the optimized parametrizations are reported in Figures 1 and 2.



Figure 1. Correlation between experimental and computed solvation free energies ΔG_{hyd} (kcal/mol, respectively, *x*-axis and *y*-axis) for a range of compounds of interest. Computed values obtained after optimization of the LJ parameters.

While the best *l* typically differs by $\Delta l = 0.1$, the average quality of all parametrizations is unchanged. In the published article,¹ the statistical measures for ΔG_{hyd} and ΔH were (RMSE = 0.36 kcal/mol, $R^2 = 0.99$) and (RMSE = 0.53 kcal/mol, $R^2 = 0.97$; see Figures 3 and 4 in ref 1), which changes to RMSE = 0.31 kcal/mol, $R^2 = 0.99$, and RMSE = 0.57 kcal/mol, $R^2 = 0.96$, using the correct script, respectively.

For one example, *N*-methyl-acetamide, the three observables $(\rho, \Delta H, \Delta G_{hyd})$ were given explicitly as a function of the scaling *l* in Table 1 of ref 1. This data has been recomputed and is reported here in Table 1. In this case the same scaling *l* = 0.95 is found to provide the best parametrization, i.e. the one with the lowest score $S = \sum_{i=1}^{3} w_i (Obs_i - Calc_i)^2$ with $w_{\rho} = 1$, $w_{\Delta H} = 3$, and $w_{\Delta G} = 5$ which differently weights the three observables.¹ The scores *S* are now larger in magnitude than in the original



Figure 2. Correlation between experimental and computed enthalpy of vaporization ΔH_{vap} (kcal/mol, respectively, *x*-axis and *y*-axis) for a range of compounds of interest. Both, MTP and LJ parameters were optimized.

Table 1. Depen	dence of ρ (g/cm^3), ΔH_{vap} ,	and ΔG_{hyd}	(both
in kcal/mol) W	hen Scaling	the Lennard-Jo	nes Param	eters ^a

scaling <i>l</i>	ρ	$\Delta H_{ m vap}$	$\Delta G_{ m hyd}$	score S		
0.9	1.13	14.24	-10.57	1.2		
0.925	1.08	13.95	-10.41	0.8		
0.95	1.00	14.11	-10.31	0.3		
0.975	0.99	13.84	-10.23	0.5		
1	0.95	13.82	-9.78	0.9		
1.025	0.92	13.68	-9.27	4.1		
1.05	0.88	13.57	-9.01	6.9		
1.075	0.84	13.29	-8.21	20.0		
1.1	0.81	13.47	-7.98	23.7		
exp	0.94 ^{2,3}	$14.2^{2,4}$	-10.08^{5}			
² In bold face is shown the value of <i>l</i> minimizing the score <i>S</i> .						

work¹ because the results from the hydration free energy simulations differ.

The current results show that the quality of the parametrizations and all conclusions from the original article remain unchanged. However, the value of the scaling l that is required for a particular quality of a parametrization changes.

AUTHOR INFORMATION

Corresponding Author

*E-mail: m.meuwly@unibas.ch.

Published: January 5, 2017

REFERENCES

(1) Hedin, F.; El Hage, K.; Meuwly, M. A Toolkit to Fit Nonbonded Parameters from and for Condensed Phase Simulations. J. Chem. Inf. Model. 2016, 56, 1479–1489.

(2) Kim, S.; Thiessen, P. A.; Bolton, E. E.; Chen, J.; Fu, G.; Gindulyte, A.; Han, L.; He, J.; He, S.; Shoemaker, B. A.; Wang, J.; Yu, B.; Zhang, J.; Bryant, S. H. PubChem Substance and Compound databases. *Nucleic Acids Res.* **2016**, *44*, D1202–D1213.

(3) CRC Handbook of Chemistry and Physics, 96th ed.; CRC: 2015.

(4) Riddick, J. A.; Bunger, W. B.; Sakano, T. K. *Physical Properties and Methods of Purification*, 4th ed.; John Wiley & Sons: New York, NY, 1985; Vol. II. Organic Solvents; p 660.

(5) Abraham, M. H.; Andonian-Haftvan, J.; Whiting, G. S.; Leo, A.; Taft, R. S. Hydrogen bonding. Part 34. The factors that influence the solubility of gases and vapours in water at 298 K, and a new method for its determination. *J. Chem. Soc., Perkin Trans.* 2 1994, 1777–1791.